than $^3Ph_2CO^*$, exactly paralleling the situation at 77 °K, leading to precisely the opposite conclusion, namely that $^3MK^*$ has an $^3n,\pi^*$ configuration in cyclohexane. In fact, in CCl_4, τ_0 for MK is even less than for Ph_2CO! Thus, no firm conclusion about the electronic configuration of $^3MK^*$ in nonpolar solvents can be made on the basis of the available evidence, and we must await further investigation into the individual decay modes (radiative, nonradiative, reversible, and irreversible reaction) which determine τ_0.

The fact that self-quenching of MK is observed is not in itself inconsistent with an n,π^* configuration of the triplet, since recent studies indicate that self-quenching is a general process for ketone triplets and is not restricted to π,π^* triplets.[9-11] Self-quenching is particularly significant kinetically in solvents such as benzene, CCl_4, and water where ketone triplet lifetimes tend to be long at room temperature, compared with lifetimes in solvents such as 2-propanol and cyclohexane. Comparison of reactivities (*e.g.*, quantum yields) of ketones in different solvents must take into account such concentration dependencies. Consequently, it is necessary to reevaluate such data[2-4] already in the literature.

In any event, these results indicate that one must be wary about making extrapolations concerning the relationship of triplet lifetimes and electronic configurations of substituted benzophenones and presumably other ketones in fluid solution at room temperature, on the basis of experimental observations in glasses at 77 °K. In addition, it would appear that τ_0 may not be sensitive to changes in electronic configuration at room temperature and thus may not be suitable as a probe for determining electronic configuration.

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[9] D. R. Kemp and P. de Mayo, *J. Chem. Soc., Chem. Commun.*, 233 (1972).

[10] D. I. Schuster and T. M. Weil, unpublished results.

[11] O. L. Chapman and G. Wampfler, *J. Amer. Chem. Soc.*, **91**, 5390 (1969); see also P. S. Engel and B. M. Monroe, *Advan. Photochem.*, **8**, 245 (1971).

Additions and Corrections

The Mechanism of the Reaction of Silver Nitrite with Alkyl Halides. The Contrasting Reactions of Silver and Alkali Metal Salts with Alkyl Halides. The Alkylation of Ambident Anions [*J. Amer. Chem. Soc.*, **77**, 6269 (1955)]. By NATHAN KORNBLUM,* ROBERT A. SMILEY, ROBERT K. BLACKWOOD, and DON C. IFFLAND, Department of Chemistry, Purdue University, Lafayette, Indiana 47907.

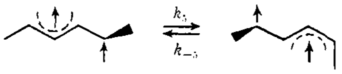
In column 1, line 38, of page 6278, "20 g of silver nitrite" read "120 g of silver nitrite."

On the Kinetics and Mechanism of the Perkow Reaction [*J. Amer. Chem. Soc.*, **94**, 1623 (1973)]. By IRVING J. BOROWITZ,* STEVEN FIRSTENBERG, GRACE B. BOROWITZ, and DAVID SCHUESSLER, Department of Chemistry, Belfer Graduate School of Science, Yeshiva University, New York, New York 10033, and Upsala College, East Orange, New Jersey 07019.

The fourth sentence of the abstract read: "ρ values for the two series are 2.37 and 1.89, respectively."

Temperature and Viscosity Effects on the Decay Characteristics of *s-trans*-1,3-Diene Triplets [*J. Amer. Chem. Soc.*, **94**, 5903 (1972)]. By JACK SALTIEL,* ALAN D. ROUSSEAU, and ALAN SYKES, Department of Chemistry, Florida State University, Tallahassee, Florida 32306.

Equation 5 is incorrect and should read



^{3}tp ^{3}cp (5)

On page 5905, line 9, read "10^{-8}" instead of "10^8." On the same page, line 14 read "faster" instead of "slower."

An Electron Spin Resonance Study of the Reaction of *tert*-Butoxy Radicals with Triphenylarsine [*J. Amer. Chem. Soc.*, **94**, 5932 (1972)]. By E. FURIMSKY, J. A. HOWARD,* and J. R. MORTON, Division of Chemistry, National Research Council of Canada, Ottawa, Canada K1 A OR9.

The second line of the ninth paragraph on page 5932 should read. "The rates were first-order in the radical concentration and independent of the triphenylarsine concentration and steady-state radical concentrations were proportional to the intensity of the initiating light."

Oxidation of Cobalt(I) Carbonyl Complexes and Cobalt(I)-Catalyzed Oxidation of Carbon Monoxide [*J. Amer. Chem. Soc.*, **94**, 6534 (1972)]. By JOHN E. BERCAW, LAI-YOONG GOH, and JACK HALPERN,* Department of Chemistry, The University of Chicago, Chicago, Illinois 60637.

The formula of **5** in Scheme 1 should read $[YCo(CN)_2(PEt_3)_2(CO_2)]^{5-}$. *Formally*, this corresponds to a CO_2 complex of cobalt(I).

Degenerate Thermal Rearrangement of Bicyclo[5.3.0^{4,8}]deca-2,5,9-triene (Lumibullvalene) [*J. Amer. Chem. Soc.*,